Chapter Seven Review

I. Identify each of the following compounds as either ionic or covalent. Name them.

C a. N_2O_3 dinitrogen trioxide

I b. $NaHCO_3$ sodium hydrogen carbonate

C c. SiO_2 silicon dioxide

I d. $Ba_3(PO_4)_2$ barium phosphate

I e. Li_2SO_4 lithium sulfate

I f. Mn_2O_3 manganese (III) oxide

C g. PBr_5 phosphorus pentabromide

I h. $Fe(MnO_4)_3$ iron (III) permanganate

C i. CS_2 carbon disulfide

I j. NaBr sodium bromide

II. Identify the following compounds as either ionic or covalent. Write the formula of the following compounds:

I a. Magnesium phosphate $Mg_3(PO_4)_2$

C b. Carbon tetrachloride CCl_4

I c. Strontium nitride Sr_3N_2

I d. Manganese (II) oxide MnO $Mn^{+2}\ O^{-2}$

I e. Iron (III) sulfite $Fe_2(SO_3)_3$

I f. Potassium oxalate $K_2C_2O_4$

I g. Aluminum Iodide AlI_3

C h. Dinitrogen pentaoxide N_2O_5

C i. Sulfur dioxide SO_2

C j. tetraphosphorus heptaoxide P_4O_7

III. Identify the following as either ionic or covalent:

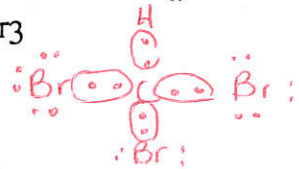
SKIP

I a. Substance has melts at 450°C; it is a crystalline structure; and when in solution it conducts electricity.

C b. Substance melts at 100°C; it is malleable and does not conduct electricity.

IV. Lewis Structures: Diagram the compound using Lewis Structures

a. $CHBr_3$



b. NaCl

c. PF_3

Name the following ionic compounds

$HgCl$	mercury (I) chloride
Ag_2SO_4	silver sulfate
$Ca(OH)_2$	calcium hydroxide
$CuSO_4$	copper (II) sulfate
Fe_2O_3	iron (III) oxide
Cu_2O	copper (I) oxide
$MgCl_2$	magnesium chloride
SnF_4	tin (IV) fluoride
$CsBr$	cesium bromide
$CaBr_2$	calcium bromide
$(NH_4)_2S$	ammonium sulfide
$PbCl_2$	lead (II) chloride
Ni_2O_3	nickel (III) oxide
$AlBr_3$	aluminum bromide
MnO_2	manganese (IV) oxide
ZnS	zinc sulfide
$BaBr_2$	barium bromide